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The energy spectrum of a system with such discrete energy levels is said to be quantized

In chemistry and atomic physics an electron shell or principal energy level may be thought of as the orbit of one or more electrons around an atom's nucleus. The configuration and energy levels of an atom selectrons determine the atom's chemical properties. Electrons are bound to the nucleus to different degrees

The outermost or valence electrons are the least tightly bound and are responsible for the formation of chemical bonds between atoms to create molecules and crystals. According to the laws of quantum mechanics a level of energy is associated with each electron configuration

In certain conditions electrons are able to move from one configuration to another by the emission or absorption of a quantum of energy in the form of a photon. Although it is sometimes stated that all the electrons in a shell have the same energy this is an approximation

However the electrons in one subshell do have exactly the same level of energy with later subshells having more energy per electron than earlier ones. In a simplistic one electron model described below the total energy of an electron is a negative inverse quadratic function of the principal quantum number leading to degenerate energy levels for each $n > 1$

[1] In more complex systems those having forces other than the nucleus-electron Coulomb force these levels split. In atomic physics and chemistry an atomic electron transition also called an atomic transition quantum jump or quantum leap is an electron changing from one energy level to another within an atom [1] or artificial atom

[2] These energy levels are discrete quantized and obtain unique energy gaps specific to a given atom. Energy levels for an electron in an atom ground state and excited states. After absorbing energy an electron may jump from the ground state to a higher energy excited state

The ground state of a quantum mechanical system is its stationary state of lowest energy. The energy of the ground state is known as the zero point energy of the system. An orbital can be occupied by a maximum of two electrons each with its own projection of spin

The simple names s orbital p orbital d orbital and f orbital refer to orbitals with angular momentum quantum number = 0 1 2 and 3 respectively. These names together with their n values are used to describe electron configurations of atoms.

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