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The use of chloroform as an incapacitating agent has become widely recognized bordering on clich through the adoption by crime fiction authors of plots involving criminals using chloroform soaked rags to render victims unconscious. Chloroform (CHCl<sub>3</sub>) is a nonflammable, clear, colourless liquid that is denser than water and has a pleasant etherlike odour.

It was first prepared in 1831.

The Scottish physician Sir James Simpson of the University of Edinburgh was the first to use it as an anesthetic in 1847. What is chloroform? Chloroform is a colorless liquid with a pleasant, nonirritating odor and a slightly sweet taste. Description: Chloroform is a colorless liquid with a pleasant, nonirritating odor and a slightly sweet taste.

It will burn only when it reaches very high temperatures.

In the past, chloroform was used as an inhaled anesthetic during surgery, but it is not used that way today. Chloroform may be released to the air from a large number of sources related to its manufacture and use, as well as its formation in the chlorination of drinking water, wastewater, and swimming pools. Aug 23 2025 With a density greater than water, chloroform is only slightly soluble in water but readily mixes with many organic solvents.

This article explores how chloroform produces its effects from its interactions at a molecular level to its historical and current applications alongside associated risks. The Molecular Mechanism of Chloroform: Chloroform is produced by mixing methane and chlorine at a temperature of 400 to 500 °C.

In this reaction, the result is a mix of chloromethane, methylene chloride, chloroform, and carbon tetrachloride.

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