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Unit 3 Worksheet1 Label the Orbital Blocks andEnergyLevels Period Orbital Name Number of Orbitals Maximum ofElectronsAssigningElectrons Examples Element Atomic Orbitals In chemistry you will need to use two equations todowith electromagnetic radiation wave speed = wavelength x frequencyenergy= Plancks constant x frequency State the units for all of the terms in both equations

Rearrange the wave speed equation to make wavelength the subject. Use the first equation to complete these wave calculations Be specific 5. What is an orbital?

6. List the orbitals that exist in the followingenergylevels

Indicate the number of these orbitals. n = 1 n = 2 n = 3 n = 4 7. Indicate the electron configuration of these elements using the shorthand notation and the orbital diagrams

For example helium 1s2 1s Elements Li C Mg Fe Kr 8 Electron Energy and LightHow does light reveal the behavior ofelectronsin an atom? Why? From fireworks to stars the color of light is useful in finding out whats in matter

The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms Explore electronenergy light emission and atomic structure with this high schoolchemistryworksheet. Covers spectra Bohr model and more The lowestenergyorbital is in Shell 1

The orbitals in shell 2 are higher inenergythan those in shell 1 but lower inenergythan those in shell 3. So as the shell number gets higher theelectronsin the orbitals have moreenergy. All atoms haveelectrons

An atom of hydrogen has only one electron whereas an atom of calcium has 20electrons Pg 117 359 Topic 1 Matter andEnergyTopic 2 The Periodic Table Topic 3 The Atomic Structure Topic 4 Chemical Bonding Topic 5 Chemical Formulas and Equations Topic 6 Moles calculations Topic 7 Solutions Topic 8 Acids Bases and Salts Topic 9 Kinetic and Equilibrium Topic 10 OrganicChemistryTopic 11 Redox and Electrochemistry Topic 12 NuclearChemistry The terms atom electrons electron orbital principleenergylevel subshell periodic table periods representative element transition elements lanthanide and actinide series s p d f Lewis valence electron dot structure valence electron and electron configuration.

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