

\$oh no! theres a negro in my wife! 4\$ Question 6f539 + Example Socratic

Calculating the concentration of excess HCL in E Socratic.

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The effect of strong base on water is to dramatically increase the concentration of OH^- ions and decrease the concentration of H_3O^+ ions. Water always contains at least small concentrations of both OH^- hydroxide and H_3O^+ hydronium ions

This is because water can react with itself in a self ionization reaction $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ At equilibrium which is attained quickly for the acid in excess is then titrated with NaOH aq of known concentration we can thus get back to the concentration or molar quantity of NaOH as it stands the question and answer are hypothetical where the combination ratio is 2:1 then 0.12 mol of nitric acid combine with 0.06 mol of Ca(OH)_2 but you have not this quantity then Ca(OH)_2 is the limiting reactant and 0.056 mol of it combines with 0.112 mol of nitric acid that you have. The inductive effect is the effect on electron density in one portion of a molecule caused by electron withdrawing or electron donating groups elsewhere in the molecule

In a covalent bond between two atoms of unequal electronegativity the more electronegative atom draws electron density towards itself. This causes the partial charges of the bond dipole

$\text{H}^+ + \text{Cl}^- \rightleftharpoons \text{HCl}$ If the $\text{pK}_a = 4.27$ $\text{pH} = 11.86$ Ethanoic acid dissociates as follows $\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$ $\text{pK}_a = 4.76$ These are equilibrium concentrations

To find the pH we need the H^+ ion concentration as $\text{pH} = -\log[\text{H}^+]$ As the base is added we get $\text{CH}_3\text{COOH} + \text{OH}^- \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_2\text{O}$ The initial moles of acid is given by $\text{pH} = -\log[\text{H}^+]$ The longer the alkyl chain attached to the hydroxyl head usually the more basic the conjugate base is and the less nucleophilic $\text{H}^+ + \text{OH}^- \rightleftharpoons \text{H}_2\text{O}$ when the acid was added to the resulting solution

The H^+ and OH^- react in a 1:1 ratio. This tells us that the number of moles of H^+ used will be equal to the number of OH^- moles in solution. Likewise 2 moles of lithium produces 2 moles of OH^- . This is also a 1:1 ratio. Since water is in excess 67.7 g MgO are needed to produce 98.0 g Mg(OH)_2

Balanced equation $\text{MgO} + \text{H}_2\text{O} \rightleftharpoons \text{Mg(OH)}_2$ Moles of magnesium hydroxide Start with the given mass of Mg(OH)_2 and convert it to moles by dividing by its molar mass 58.319 g/mol

Since molar mass is a fraction g/mol we can divide by multiplying by the reciprocal of the molar mass mol/g $\text{pH} = 14 - \log[\text{H}^+]$ $\text{pH} = 14 - \log(0.024462 \text{M})$ $\text{pH} = 11.86$ We can find the concentration of H^+ or H_3O^+ by three ways One is by the ICE table but this is a 5% rule and the other is square root which is absolutely correct and the other is Ostwald's law of dilution Let's set up an ICE table

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