#oh thats tight 4# Question 6f539 + Example Socratic Calculating the concentration of excess HCL in E Socratic.

The effect of strong base on water is to dramatically increase the concentration of OH^- ions and decrease the concentration of H_3O^+ ions. Water always contains at least small concentrations of both OH^- hydroxide and H_3O^+ hydronium ions

This is because water can react with itself in a self ionization reaction $2 H_2O \rightleftharpoons H_3O^+ + OH^-$ At equilibrium which is attained quickly for The acid in excess is then titrated with $NaOH$ aq of KNOWN concentration we can thus get back to the concentration or molar quantity of $M OH 2$ as it stands the question and answer are hypothetical where the combination ratio is 2 1 then 0 12 mol of nitric acid combine with 0 06 mol of $Ca OH_2$ but you have not this quantity then $Ca OH_2$ is the limitant reactant and 0 056 mol of it combines with 0 112 mol of nitric acid thatyou have The inductive effect is the effect on electron density in one portion of a molecule caused by electron withdrawing or electron donating groups elsewhere in the molecule

In a covalent bond between two atoms of unequal electronegativity the more electronegative atom draws electron density towards itself. This causes the and charges of the bond dipole

H^+ + Cl^- If the 4.27 11.86 Ethanoic acid dissociates sf $CH_3COOH \rightleftharpoons CH_3COO^- + H^+$ sf $K_a= \frac{[CH_3COO^-][H^+]}{[CH_3COOH]} = 1.8 \times 10^{-5}$ These are equilibrium concentrations

To find the pH we need the sf H^+ ion concentration sf $[H^+] = K_a \times \frac{[CH_3COOH]}{[CH_3COO^-]}$ As the base is added we get sf $CH_3COOH + OH^- \rightarrow CH_3COO^- + H_2O$ The initial moles of acid is given by The longer the alkyl chain attached to the hydroxyl head usually the more basic the conjugate base is and the less nucleophilic $H^+ + OH^- \rightarrow H_2O$ when the acid was added to the resulting solution

The H^+ and OH^- react in a 1 1 ratio. This tells us that the number of moles of H^+ used will be equal to the number of OH^- moles in solution. Likewise 2 moles of lithium produces 2 moles of OH^-. This is also a 1 1 ratio Since water is in excess 67.7 g MgO are needed to produce 98.0 g $Mg OH_2$

Balanced equation $MgO s + H_2O l \rightarrow Mg OH_2 s$ Moles magnesium hydroxide Start with the given mass of $Mg OH_2$ and convert it to moles by dividing by its molar mass 58.319 g mol

Since molar mass is a fraction g mol we can divide by multiplying by the reciprocal of the molar mass mol g pH = 1.61151 $OH^- = 4.08797 \times 10^{13} M$ HF = 0.855538M $H^+ = 0.024462M$ $F^- = 0.024462M$ $HF + H_2O = H_3O^+ + F^-$ We can find the concentration of H^+ or H_3O^+ by three ways One is by the ICE table but this is a 5% rule and the other is square root which is absolutely correct and the other is Ostwald s law of dillution Let s set up an ICE table

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