@oh thats tight@ Question 71ce2 Socratic Question 6f539 + Example Socratic Calculating the concentration of excess.

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H^+ +OH^ > H_2O when the acid was added to the resulting solution. The H^+ andOH^ react in a 1 1 ratio. This tells usthatthe number of moles of H^+ used will be equal to the number ofOH^ moles in solution. Likewise 2 moles of lithium produces 2 moles ofOH^

This is also a 1 1 ratio The effect of strong base on water is to dramatically increase the concentration ofOH^ ions and decrease the concentration of H_3O^+ ions. Water always contains at least small concentrations of bothOH^ hydroxide and H_3O^+ hydronium ions

This is because water can react with itself in a self ionization reaction 2 H_2O harr H_3O^+ +OH^ At equilibrium which is attained quickly for The acid in excess is then titrated with N aOH aq of KNOWN concentration we can thus get back to the concentration or molar quantity of M OH 2 as it stands the question and answer are hypothetical where the combination ratio is 2 1 then 0 12 mol of nitric acid combine with 0 06 mol of Ca OH _2 but you have not this quantity then Ca OH _2 is the limitant reactant and 0 056 mol of it combines with 0 112 mol of nitric acidthatyou have The inductive effect is the effect on electron density in one portion of a molecule caused by electron withdrawing or electron donating groups elsewhere in the molecule

In a covalent bond between two atoms of unequal electronegativity the more electronegative atom draws electron density towards itself. This causes the and charges of the bond dipole

H ^ + Cl ^ If the 4.27 11.86 Ethanoic acid dissociates sf CH_3COOHrightleftharpoonsCH_3COO^ +H^+ sf K_a= [CH_3COO^] [H^+] [CH_3COOH] =1.8xx10^ 5 These are equilibrium concentrations

To find the pH we need the sf H^+ ion concentration sf [H^+]=K_axx [CH_3COOH] [CH_3COO^] As the base is added we get sf CH_3COOH+OH^ rarrCH_3COO^ +H_2O The initial moles of acid is given by The longer the alkyl chain attached to the hydroxyl head usually the more basic the conjugate base is and the less nucleophilic Since water is in excess 67.7 g MgO are needed to produce 98.0 g Mg OH _2

Balanced equation MgO s + H _2 O l rarr Mg OH _2 s Moles magnesium hydroxide Start with the given mass of Mg OH _2 and convert it to moles by dividing by its molar mass 58.319 g mol

Since molar mass is a fraction g mol we can divide by multiplying by the reciprocal of the molar mass mol g pH = 1.61151 OH^ = 4.08797 * 10 ^ 13M HF = 0.855538M H^+ = 0.024462M F^ = 0.024462M HF + H_2O = H_3O^+ + F^ We can find the concentration of H^+ or H_3O^+ by three ways One is by the ICE table but this is a 5% rule and the other is square root which is absolutely correct and the other

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