

\$oh! zone\$ Question 71ce2 Socratic Question 6f539 + Example Socratic Calculating the concentration of excess.

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$H^+ + OH^- \rightarrow H_2O$ when the acid was added to the resulting solution. The H^+ and OH^- react in a 1:1 ratio. This tells us that the number of moles of H^+ used will be equal to the number of OH^- moles in solution. Likewise 2 moles of lithium produces 2 moles of OH^-

This is also a 1:1 ratio. The effect of strong base on water is to dramatically increase the concentration of OH^- ions and decrease the concentration of H_3O^+ ions. Water always contains at least small concentrations of both OH^- hydroxide and H_3O^+ hydronium ions.

This is because water can react with itself in a self ionization reaction $2 H_2O \rightleftharpoons H_3O^+ + OH^-$. At equilibrium which is attained quickly for the acid in excess is then titrated with $N aOH$ aq of KNOWN concentration we can thus get back to the concentration or molar quantity of $M OH_2$ as it stands the question and answer are hypothetical where the combination ratio is 2:1 then 0.12 mol of nitric acid combine with 0.06 mol of $Ca(OH)_2$ but you have not this quantity then $Ca(OH)_2$ is the limitant reactant and 0.056 mol of it combines with 0.112 mol of nitric acid that you have. The inductive effect is the effect on electron density in one portion of a molecule caused by electron withdrawing or electron donating groups elsewhere in the molecule.

In a covalent bond between two atoms of unequal electronegativity the more electronegative atom draws electron density towards itself. This causes the partial charges of the bond dipole

$H^- + Cl^- \rightleftharpoons HCl$ If the $K_a = 4.27 \times 10^{-5}$ Ethanoic acid dissociates as $CH_3COOH \rightleftharpoons CH_3COO^- + H^+$ at $K_a = 1.8 \times 10^{-5}$ These are equilibrium concentrations

To find the pH we need the H^+ ion concentration as $[H^+] = K_a \times [CH_3COOH] / [CH_3COO^-]$. As the base is added we get $CH_3COOH + OH^- \rightleftharpoons CH_3COO^- + H_2O$. The initial moles of acid is given by M . The longer the alkyl chain attached to the hydroxyl head usually the more basic the conjugate base is and the less nucleophilic. Since water is in excess 67.7 g MgO are needed to produce 98.0 g $Mg(OH)_2$.

Balanced equation $MgO(s) + H_2O(l) \rightarrow Mg(OH)_2(s)$ Moles magnesium hydroxide Start with the given mass of $Mg(OH)_2$ and convert it to moles by dividing by its molar mass 58.319 g/mol

Since molar mass is a fraction g/mol we can divide by multiplying by the reciprocal of the molar mass mol/g $pH = 1.61151$ $OH^- = 4.08797 \times 10^{-13} M$ $HF = 0.855538 M$ $H^+ = 0.024462 M$ $F^- = 0.024462 M$ $HF + H_2O = H_3O^+ + F^-$ We can find the concentration of H^+ or H_3O^+ by three ways. One is by the ICE table but this is a 5% rule and the other is square root which is absolutely correct and the other

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